**Making Cents of Penny Composition**

**Part II**

**Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

Directions:

1. Review the compositions listed below:

In 1962, the cent's tin content, which was quite small to begin with, was removed. That made the metal composition of the cent 95% copper and 5% zinc.

The alloy remained the same until 1982, when the composition was changed to 97.5% zinc and 2.5% copper (copper plated zinc). Cents of both compositions appeared in that year.

1. Use your data compiled in the Data Sheet for Part I to determine the mass and number of moles of each element (copper and zinc) in each one of your pennies.
2. Enter your results in the table below:

**Data Sheet for Part II**

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Penny** | **Date** | **Mass of Cu** | **Mass of Zn** | **Moles of Cu** | **Moles of Zn** |
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